

Hz), 161.1, 162.0, 201.2 ppm. Anal. Calcd. (%) for $C_{24}H_{16}O_4Cl$: C, 68.17; H, 3.81. Found: C, 68.24; H, 3.87.

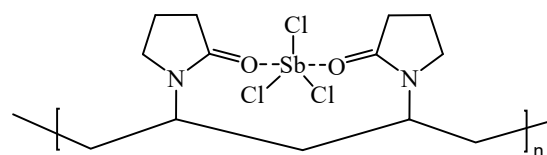
3- (3- (4- chlorophenyl)- 1- (4- (dimethylamino)phenyl)- 3-oxopropyl)-4-hydroxy-2H-chromen-2-one (**3j**):

White solid. m.p.= 176-178 °C. FT-IR (KBr): $\bar{\nu}$ = 756, 815, 1095, 1207, 1398, 1456, 1517, 1564, 1618, 1654, 1699, 2854, 2920, 3436 cm^{-1} . 1H NMR (400 MHz, $CDCl_3$): δ = 2.96 (s, 6H), 3.70 (dd, J = 19.0, 2.8 Hz, 1H), 4.35 (dd, J = 19.0, 9.6 Hz, 1H), 4.87 (dd, J = 9.6, 2.8 Hz, 1H), 6.72 (d, J = 8.8 Hz, 2H), 7.24-7.31 (m, 4H), 7.48-7.53 (m, 3H), 7.96 (dd, J = 8.0, 1.2 Hz, 1H), 8.40 (d, J = 8.8 Hz, 2H), 9.33 (s, 1H) ppm. ^{13}C NMR (100 MHz, $CDCl_3$): δ = 31.9, 34.6, 40.5, 40.6, 108.1, 112.5, 116.2, 116.7, 123.7, 123.8, 128.6, 128.8, 129.1, 130.0, 131.6, 134.3, 140.7, 149.4, 152.8, 160.6, 162.9, 201.6 ppm. Anal. Calcd. (%) for $C_{26}H_{22}NO_4Cl$: C, 69.72; H, 4.95, N, 3.13. Found: C, 70.02; H, 5.01, N, 3.17.

3. Results and Discussion

In this study, PVPP-SbCl₃ is obtained by a simple reaction of antimony(III) chloride with the polyvinylpyrrolidone (Scheme 2).

Characterization of the polyvinylpyrrolidone supported antimony(III) chloride was performed by recording the Fourier transform infrared spectroscopy (FT-IR) spectrum of PVPP-SbCl₃ (Fig. 1). For polyvinylpyrrolidone, the C-N stretch band, CH₂ bending band, the C=O stretch band, and the CH stretch band were found to be 1288 cm^{-1} , 1429 cm^{-1} , 1649 cm^{-1} , and 2956 cm^{-1} respectively (Fig. 1a). According to FT-IR analysis of PVPP-SbCl₃ indicated in Fig. 1b, the stretching vibrations of C=O for PVPP-SbCl₃ was



Scheme 2. The proposed chemical structure of PVPP-SbCl₃ complex.

observed at 1637 cm^{-1} . The good complexation of carbonyl functional groups of PVPP with 12 cm^{-1} red shift was observed for PVPP-SbCl₃. These results provided the evidences that antimony(III) chloride was successfully attached to the polyvinylpyrrolidone.

The thermal behavior of PVPP-SbCl₃ was shown in Fig. 2. The thermal analysis of PVPP-SbCl₃ showed two main decreasing peaks. First peak appears at temperature around 220–280 °C due to the loss of the antimony(III) chloride group from the polymer surface. This is followed by a second peak at 320–370 °C, corresponding to the decomposition of the pyrrolidone functional group. The thermal analysis data showed that the catalyst is stable up to 200 °C.

Surface morphological study of PVPP-SbCl₃ was explored by scanning electron microscope (Fig. 3). The results of SEM analysis revealed that the SbCl₃ was properly dispersed within the PVPP matrix, which was directly related to good complexation between antimony(III) chloride and PVPP functional groups.

The elemental composition of the SEM image of PVPP-SbCl₃ is presented in Fig. 4. The EDX spectrum shows the presence of C, N, O, Sb and Cl in the catalyst. According to the EDX analysis, the antimony content in the catalyst was obtained 10.40 w%.

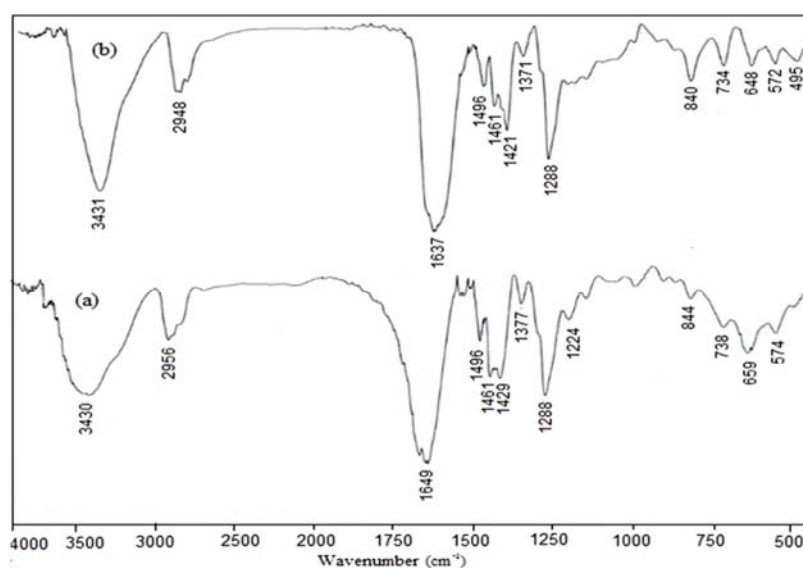


Fig. 1. FT-IR spectra of (a) PVPP, (b) PVPP-SbCl₃.

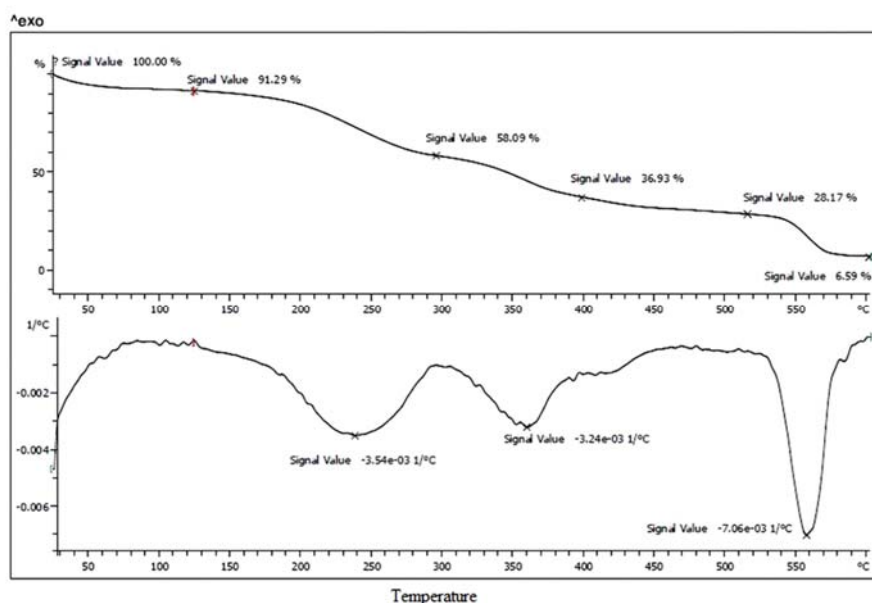


Fig. 2. The thermal analysis diagram of PVPP-SbCl₃.

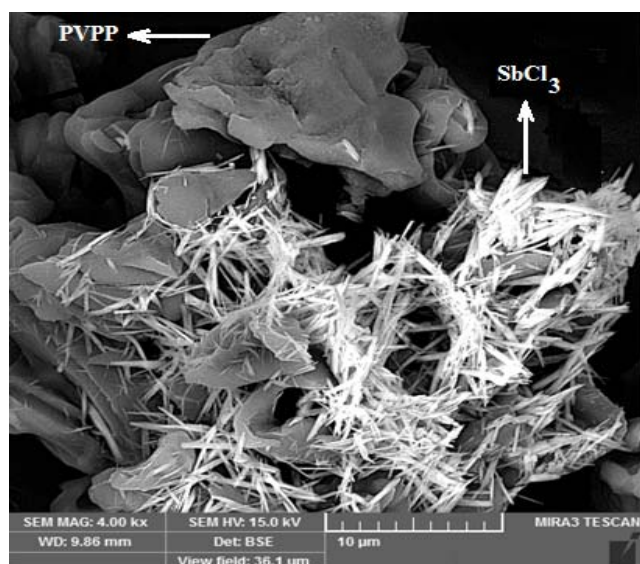


Fig. 3. SEM image of PVPP-SbCl₃.

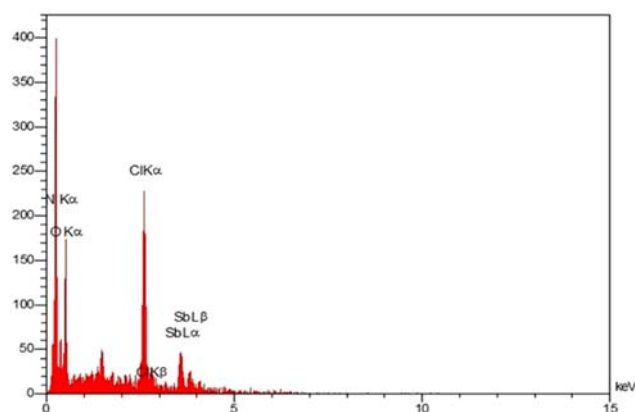


Fig. 4. The EDX spectrum of PVPP-SbCl₃.

In order to optimize the reaction conditions and get the best catalytic activity, the reaction of chalcone and 4-hydroxycoumarin was examined as a model reaction in the several catalysts. In this study, it was observed that polyvinylpyrrolidone-SbCl₃ in chloroform under reflux conditions is more efficient with respect to the reaction time and yields of the desired product (Table 1).

Furthermore, the reaction of chalcone and 4-hydroxycoumarin was examined in several solvents. In this investigation, it was perceived that SbCl₃ immobilized on polyvinylpyrrolidone in chloroform under reflux conditions is more efficient with respect to the reaction time and efficiency of the desired product (Table 2).

Table 1. Synthesis of 3a in different catalysts.^a

Entry	Catalyst	Amount (g)	Yield (%) ^b
1	nano-Fe ₃ O ₄	0.057	70
2	nano-ZnFe ₂ O ₄	0.06	65
3	nano-SnO ₂	0.037	50
4	nano-MgO	0.01	42
5	nano-ZnO	0.02	45
6	SbCl ₃	0.05	74
7	PVPP-SbCl ₃	0.05	85

^aReaction conditions: Chalcone (1 mmol), 4-hydroxycoumarin (1 mmol) in CHCl₃ (5 mL) at reflux conditions after 4h.

^bIsolated yields.

Table 2. Synthesis of **3a** in different solvents.^a

Entry	Solvent	Yield (%) ^b
1	Solvent-free	30
2	n-hexane	20
3	EtOH (96%)	80
4	MeOH	60
5	EtOAc	40
6	CHCl ₃	85

^aReaction conditions: Chalcone (1 mmol), 4-hydroxycoumarin (1 mmol), PVPP-SbCl₃ (0.05 g) at reflux conditions after 4h.

^bIsolated yields.

After optimizing the reaction conditions, a diversity of chromenylphenylpropanones was synthesized using PVPP-SbCl₃ in good yields (Table 3, entries 1-10). The reactions worked well with all chalcones bearing electron-donating or electron-withdrawing substituent.

To check the reusability of the catalyst, it was employed in the synthesis of **3a**, four cycles under the optimum conditions. The catalyst powder was recovered by easy filtration and washed with dichloromethane. Then, according to the amount of catalyst, the required amount of fresh chalcone and 4-hydroxycoumarin were added. The results showed that the catalyst can be reused four consecutive times without significant loss of its catalytic activity (The yields were 85, 82, 80 and 78%, respectively).

In order to examine the efficiency of the present method for the synthesis of chromenylphenyl propanones, compound **3a** was compared with some of those reported in the literature (Table 4). Our results can be compared with previously reported data when all terms, including yields, reaction times, and reaction conditions are taken into account.

Table 3. Synthesis of chromenylphenylpropanones catalyzed by PVPP-SbCl₃.^a

R	R'	Product	Time (h)	Yields (%) ^b	m.p. (°C)		Ref.
					Found	Reported	
H	H	3a	4	85	162-164	160-161	[8]
H	4-Cl	3b	3.5	83	177-179	179-181	[8]
4-Cl	H	3c	3.5	84	180-182	183-184	[10]
4-Br	H	3d	3.5	85	194-196	196-198	[10]
H	4-MeO	3e	4	74	176-177	175-178	[8]
3-NO ₂	H	3f	3.5	82	187-189	----	
4-Cl	4-Cl	3g	3.5	86	154-156	156-157	[11]
4-MeO	H	3h	4	70	124-126	125-126	[9]
4-F	4-Cl	3i	3.5	84	184-186	----	
4-(Me) ₂ N	4-Cl	3j	4	75	176-178	----	

^aReaction and condition: Chalcones (1 mmol), 4-hydroxycoumarin (1 mmol), and PVPP-SbCl₃ (0.05 g) in chloroform (5 mL) at reflux conditions.

^bAll yields refer to isolated products.

Table 4. Comparison of PVPP-SbCl₃ with some other catalysts for synthesis of **3a**.

Entry	Catalyst	Reaction conditions	Time/h	Yield (%)	Ref.
1	-	H ₂ O/ reflux	57	26	[8]
2	4-pyrrolidinopyridine	CHCl ₃ / reflux	24	62	[9]
3	-	EtOH/reflux	14	48	[11]
4	PVPP-SbCl ₃	CHCl ₃ / reflux	4	85	This work

4. Conclusions

In this study, we have extended a protocol for the synthesis of chromenylphenylpropanon derivatives as warfarine-like analogues by the reaction of 4-hydroxycoumarin with α,β -unsaturated compounds in the presence of polyvinylpyrrolidone supported antimony(III) chloride (PVPP-SbCl₃) as a reusable new polymeric Lewis acid catalyst in chloroform at reflux conditions.

Acknowledgements

Financial support by Rasht Branch, Islamic Azad University Grant No. 4.5830 is gratefully acknowledged.

References

- [1] R.D.H. Murray, J. Méndez, S.A. Brown, The natural coumarins: Occurrence, chemistry and biochemistry, Wiley, New York, 1982.
- [2] D. Egan, R. O'Kennedy, E. Moran, D. Cox, E. Prosser, R.D. Thornes, Drug Metab. Rev. 22 (1990) 503-529.
- [3] K.C. Fylaktakidou, D.J. Hadjipavlou-Litina, K.E. Litinas, D.N. Nicolaidis, Curr. Pharm. 10 (2004) 3813-3833.
- [4] F. Dall'Acqua, W.M. Horspool, D. Vedaldi, P.S. Song, Handbook of organic photochemistry and photobiology, CRC Press, Boca Raton, 1995.
- [5] A. Guiotto, A. Chilini, P. Manzini, F. Dall'Acqua, F. Bordin, P. Rodighiero, Farmaco 50 (1995) 479-488.
- [6] C. Mahidol, P. Ploypradith, P. Sahakitpichan, S. Wongbundit, S. Ruchirawat, Angew. Chem. Int. Ed. 43 (2004) 866-868.
- [7] X. Li, Y. Zhao, T. Wang, M. Shi, F. Wu, Dyes Pigm. 74 (2007) 108-112.
- [8] I. Manolov, N.D. Danchev, Arch. Pharm. 336 (2003) 83-94.
- [9] O. Talhi, J.A. Fernandes, D.C.G.A. Pinto, F.A. Almeida Paz, A.M.S. Silva, J. Mol. Struct. 1094 (2015) 13-21.
- [10] S. Sabir, N. Rashid, S. Naz, B. Masood, Int. J. Pharm. Pharm. Sci. 5 (2013) 177-181.
- [11] A. A. Shkell, O.A. Mazhukina, O.V. Fedotova, Chem. Heterocycl. Compd. 47 (2011) 656-657.
- [12] Y.H. Liu, Z.H. Zhang, T.S. Li, Synthesis (2008) 3314-3318.
- [13] Z.H. Zhang, Y.H. Liu, Catal. Commun. 9 (2008) 1715-1719.
- [14] G. Maiti, P. Kundu, Tetrahedron Lett. 47 (2006) 5733-5736.
- [15] D. Mahajan, B.A. Ganai, R.L. Sharma, K.K. Kapoor, Tetrahedron Lett. 47 (2006) 7919-7921.
- [16] K.K. Kapoor, B.A. Ganai, S. Kumar, C.S. Andotra, Can. J. Chem. 84 (2006) 433-437.
- [17] A. Srinivasa, B.P. Nandeshwarappa, B.M. Kiran, K.M. Mahadevan, Phosphorus Sulfur Silicon Relat. Elem. 182 (2007) 2243-2249.
- [18] G. Aiti, P. Kundu, Synth. Commun. 37 (2007) 2309-2316.
- [19] I. Capanec, M. Litvić, M. Filipan-Litvić, I. Grüngold, Tetrahedron 63 (2007) 11822-11827.
- [20] G. Maiti, R.N. Bhattacharya, R. Karamkar, Indian J. Chem. 51B (2012) 302-307.
- [21] J. Lu, P.H. Toy, Chem. Rev. 109 (2009) 815-838.
- [22] M. Yosefzadeh, M. Mokhtary, Iran. J. Catal. 6 (2016) 153-159.
- [23] M. Mokhtary, S. Refahati, Dyes Pigm. 99 (2013) 378-381.
- [24] M. Mokhtary, S.A. Mirfarjood Langroudi, Monatsh. Chem. 145 (2014) 1489-1494.
- [25] M. Goudarzvand Chegini, M. Mokhtary, Polycycl. Aromat. Compd. 37 (2017) 63-72.